

Behavior Of Gases Test B Answers

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1. Gas particles are so small in relation to the distances between them that their individual volumes can be assumed to be insignificant. 2. No attractive or repulsive forces exist between gas particles. 3. Gas particles are in constant random motion, traveling in independent straight paths until they collide with each other or other objects.

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The gas laws describe the relationships among the gas properties of pressure, temperature, and A) Mass B) Shape C) energy D) Volume D) Volume If you increase the temperature and pressure of a gas in a sealed container, particles of the gas will A) have more energy B) move more quickly C) exert greater pressure D) all of the above

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1. Gas particles don't attract one another. 2. Gas particulars are in random, constant, straight line motion. 3. When particles collide with each other (or any container), the collisions are said to be elastic. 4. The volume that gas particles take up is negligible. And from the distance between particles is relatively great.

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a. There is an increase in the number of collisions between particles and the walls of the container. b. There is an increase in the temperature of the gas. c. There is a decrease in the volume of the gas. d. There is an increase in the force of the collisions between the particles and the walls of the container.

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The particle theory is used to explain the properties of solids, liquids and gases. The strength of bonds (attractive forces) between particles is different in all three states.

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