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required when 100% of the hexadecane is converted according to the stoichiometric equation. (I) From the stoichiometry, production of 27.3 gmol cells requires $27.3/1.65 = 16.5$ gmol = 16.5 gmol \times 226.4 g gmol⁻¹ = 3736 g = 3.74 kg hexadecane. Answer: 3.74 kg.

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?1 Answer: $L^2 T^{-1}$ 2.10 Dimensional homogeneity and gc

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